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by See Notes Multiple Contributors
CH3COO– (which is the conjugate acid of acetic acid) can be written as $\text{H}_3\text{COO}^-$. The dissociation of water can be represented as $\text{H}_2\text{O} + \text{H}^+ \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$. In order to calculate the concentration of $\text{H}_3\text{O}^+$, and hence the pH, we must also remember that the volume of the solution and the concentration of the acid are important factors. The volume of the solution must be known to calculate the number of moles of acid used. The concentration of the acid can be found by dividing the number of moles of acid by the volume of the solution.

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